Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Q: What is the name of the assessment for which I am studying?

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| **Unit 7: The Periodic Table of Elements** |

Q: When is the assessment?

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| **On Google Classroom** |

Q: What concepts should I study?

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| * The History of the Periodic Table – How did Mendeleev organize the elements? How is it organized today? * The types of elements on the PT and their properties (metals, nonmetals, metalloids) * How valence electrons behave in metals, nonmetals, metalloids and noble gases (do they gain e-, lose e-, do both or do nothing?) * What are the phases (solids, liquids or gases) of the elements at STP? * How nuclear charge changes across a period and down a group * How electronegativity changes across a period and down a group * How ionization energy changes across a period and down a group * How atomic radius changes across a period and down a group * How ionic radius differs from atomic radius for metals and nonmetals * WHY trends occur in atomic radius, electronegativity, ionization energy and reactivity across a period and down a group * The LOCATIONS of the best metal and best nonmetal on the PT and WHY * Characteristics of alkali metals, alkaline earth metals, transition metals, halogens and noble gases * The relationship between group number and valence electrons * The relationship between period number and energy shells |

Q: What examples can I use to study?

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| Unit 7 – The Periodic Table Note Packet  **Practice Packet/Tests (Answer keys on classroom)**  Periodic Table Worksheets (Classwork/Homework)  Orange Review Book – Topic #5  Create your own Castle Learning (click the “Periodic Table” box for questions)  Chemistry Reference Tables – Table S, Periodic Table of Elements |