Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Q: What is the name of the assessment for which I am studying?

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| **Unit 7: The Periodic Table of Elements** |

Q: When is the assessment?

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| **On Google Classroom** |

Q: What concepts should I study?

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| * The History of the Periodic Table – How did Mendeleev organize the elements? How is it organized today?
* The types of elements on the PT and their properties (metals, nonmetals, metalloids)
* How valence electrons behave in metals, nonmetals, metalloids and noble gases (do they gain e-, lose e-, do both or do nothing?)
* What are the phases (solids, liquids or gases) of the elements at STP?
* How nuclear charge changes across a period and down a group
* How electronegativity changes across a period and down a group
* How ionization energy changes across a period and down a group
* How atomic radius changes across a period and down a group
* How ionic radius differs from atomic radius for metals and nonmetals
* WHY trends occur in atomic radius, electronegativity, ionization energy and reactivity across a period and down a group
* The LOCATIONS of the best metal and best nonmetal on the PT and WHY
* Characteristics of alkali metals, alkaline earth metals, transition metals, halogens and noble gases
* The relationship between group number and valence electrons
* The relationship between period number and energy shells
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Q: What examples can I use to study?

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| Unit 7 – The Periodic Table Note Packet **Practice Packet/Tests (Answer keys on classroom)**Periodic Table Worksheets (Classwork/Homework)Orange Review Book – Topic #5 Create your own Castle Learning (click the “Periodic Table” box for questions)Chemistry Reference Tables – Table S, Periodic Table of Elements |