

name:

date:

Periodic Table Exam

1. ____ B
2. ____ A
3. ____ D
4. ____ A
5. ____ A
6. ____ A
7. ____ A
8. ____ D
9. ____ C
10. ____ C
11. ____ C
12. ____ B
13. Bi
14. Period 3
15. A) Noble Gases
 - b) They have a full valence shell of electrons.
16. a) Size of an atom
 - b) Radius increases
 - c) # of electron shells (energy levels) increases
 - d) Radius decreases
 - e) More protons, so stronger nuclear pull, so size gets smaller (shells are sucked in)
17. a) How strongly an atom holds onto its electrons (how much energy it takes to remove them)
 - b) It decreases

c) The more shells, the further the valence e⁻ are from the nucleus, so attraction is less and they are easier to remove (require less energy to remove)

18. It only has 1 valence electron (very unstable because not full) and low ionization energy because weak nuclear pull and big amount of shells so easy to lose it.