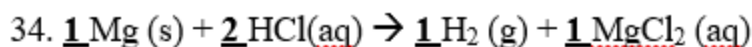


## Review Packet Answer Key

### Redox (Topic 9 in your review book)

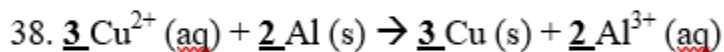
- |       |       |
|-------|-------|
| 1. 1  | 17. 2 |
| 2. 1  | 18. 1 |
| 3. 1  | 19. 3 |
| 4. 4  | 20. 4 |
| 5. 4  | 21. 3 |
| 6. 4  | 22. 3 |
| 7. 2  | 23. 1 |
| 8. 4  | 24. 3 |
| 9. 2  | 25. 3 |
| 10. 4 | 26. 4 |
| 11. 4 | 27. 2 |
| 12. 4 | 28. 4 |
| 13. 4 | 29. 3 |
| 14. 4 | 30. 2 |
| 15. 1 | 31. 4 |
| 16. 4 | 32. 4 |

33. **Mg is more reactive** and correctly oxidizes, while **H<sup>+</sup> is less reactive** and correctly reduces.



36. The cell is connected to a **battery**

37. From **-2 to 0**.

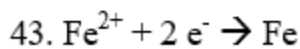


39. The Al(s) is oxidized, and therefore **loses electrons** to form Al<sup>3+</sup> (aq) in solution.

40. The salt bridge permits the **migration of ions** between the cells (and helps to close the circuit!)

41. **Oxidation numbers are changing** from the reactant to product side (Cu<sup>2+</sup> → Cu; Fe → Fe<sup>2+</sup>)

42. Electrons



44. Zn is the most reactive, followed by Fe, then Cu being the least reactive.